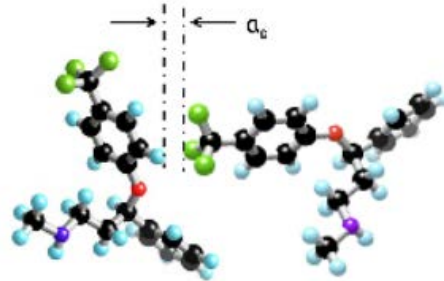


## QUIZ on Lecture P1\_Wk1\_L4

1. It's a good idea to have a distance in mind which specifies the smallest separation possible between two small neutral molecules. Call this "touching distance" distance  $a_0$ . A reasonable estimate for  $a_0$  might be

- a) ~0.05 nm
- b) ~0.3 nm
- c) ~1.5 nm
- d) ~5.0 nm



2. Suppose you have a small atomic system in thermal equilibrium with a large reservoir at temperature  $T$ . Assume the small system can have three discrete energies  $E_1$ ,  $E_2$ , and  $E_3$ . Also assume there is only one way of populating each energy level (ie. All the energy states have equal weight). The thermal average of  $E$  might best be described by

$$d) \langle E \rangle = \frac{E_1 e^{-E_1/kBT} + E_2 e^{-E_2/kBT} + E_3 e^{-E_3/kBT}}{e^{-E_1/kBT} + e^{-E_2/kBT} + e^{-E_3/kBT}}$$

3. A molecule with a permanent dipole moment

- a) cannot electrostatically interact with a non-polar molecule
- b) can interact with the dipole moment it induces in a nearby non-polar molecule
- c) requires an electric field produced by a nearby ion before it can interact with a non-polar molecule
- d) can only interact with the non-polar molecule at temperatures near 0K

4. If an ion interacts with a molecule having a permanent electric dipole moment that is thermally free to rotate, the electrostatic interaction energy

- a) is constant, independent of temperature
- b) increases linearly with increasing temperature
- c) is always zero
- d) decreases with increasing temperature as  $1/T$

5. Suppose under certain circumstances, electric fields between molecules can approach  $5 \times 10^9 \text{ V/m}$ . Estimate the induced dipole moment (in Debye) for an argon atom with an electronic polarizability volume of  $1.66 \times 10^{-30} \text{ m}^3$  if the atom encounters such an intermolecular electric field.

- a) 0.10 D
- b) 0.28 D
- c) 0.68 D
- d) 1.62 D

Solution:

$$\begin{aligned} p &= \alpha_e E = (4\pi\epsilon_0 a)E \\ &= 4\pi (8.85 \times 10^{-12} \text{ C}^2/\text{Nm}^2) (1.66 \times 10^{-30} \text{ m}^3) (5 \times 10^9 \text{ V/m}) \\ &= 9.23 \times 10^{-31} \text{ Cm} \times \frac{1\text{D}}{3.33 \times 10^{-30} \text{ Cm}} = 0.28\text{D} \\ \text{units: } &\frac{\text{C}^2}{\text{Nm}^2} \times \text{m}^3 \times \text{V/m} = \frac{\text{C}^2 \text{m}^3 \text{J/C}}{\text{Nm}^3} = \frac{\text{CNm}}{\text{N}} = \text{Cm} \end{aligned}$$