

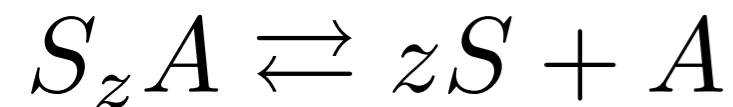
Introduction to the Materials Science of

Rechargeable Batteries

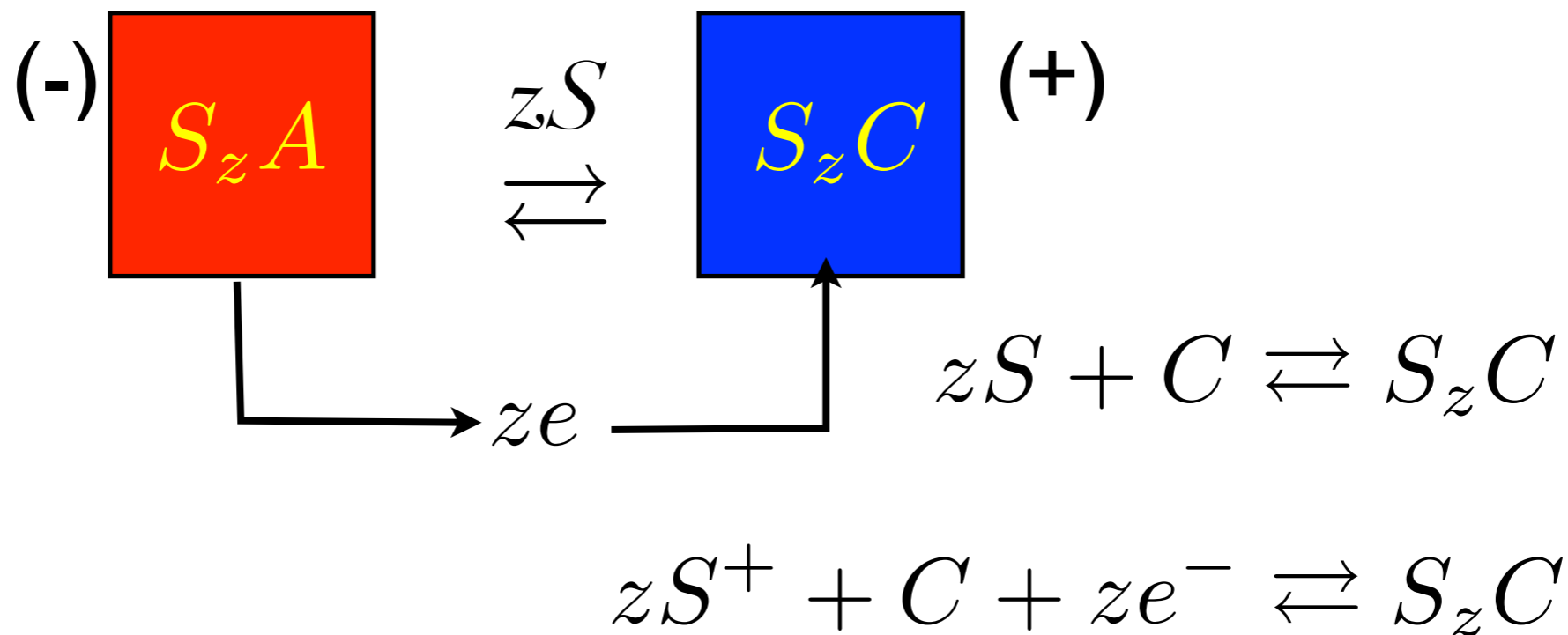
Week 1: Basic Concepts, Fundamentals, and Definitions
Lecture 1.2: Charge Figures of Merit in a Battery

By R. Edwin Garcia
Associate Professor of Materials Engineering
Purdue University

Charge in a Battery



(electrolyte)



C_T^C : cathode solubility limit

Battery Charge

For Each Phase

$$Q = z \times C_T \times V \times eN_a$$

valence

solubility
limit

volume of
electrode

electronic charge

Avogadro's Number

$$F = eN_a$$

Faraday's constant

For the Device

$$Q = A \int_{t_o}^{t_f} I dt$$

Units:

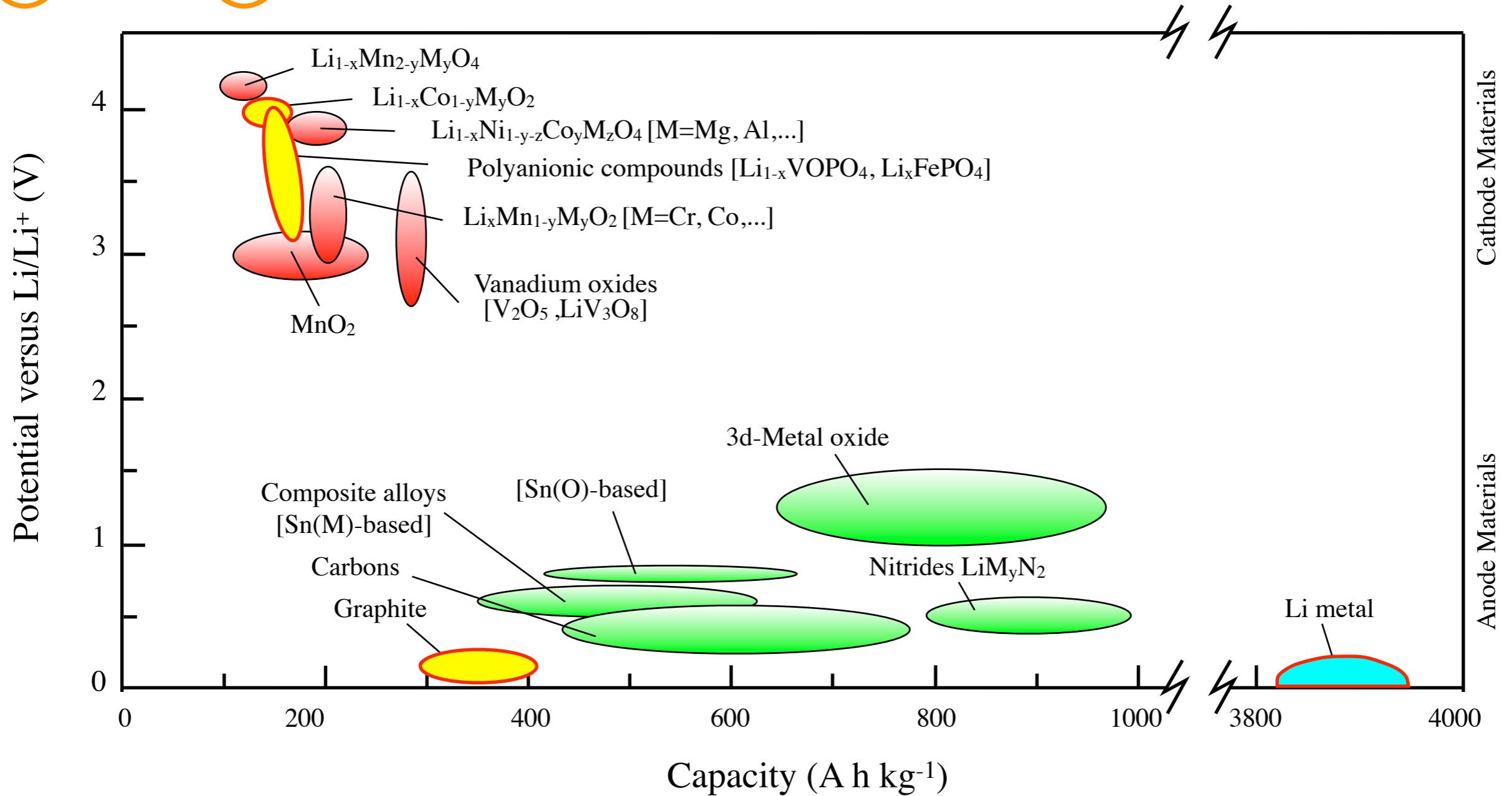
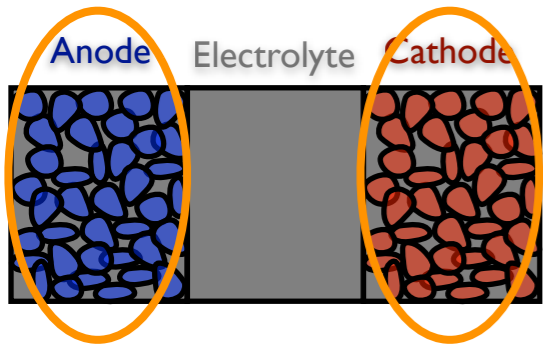
mA h

Wish List

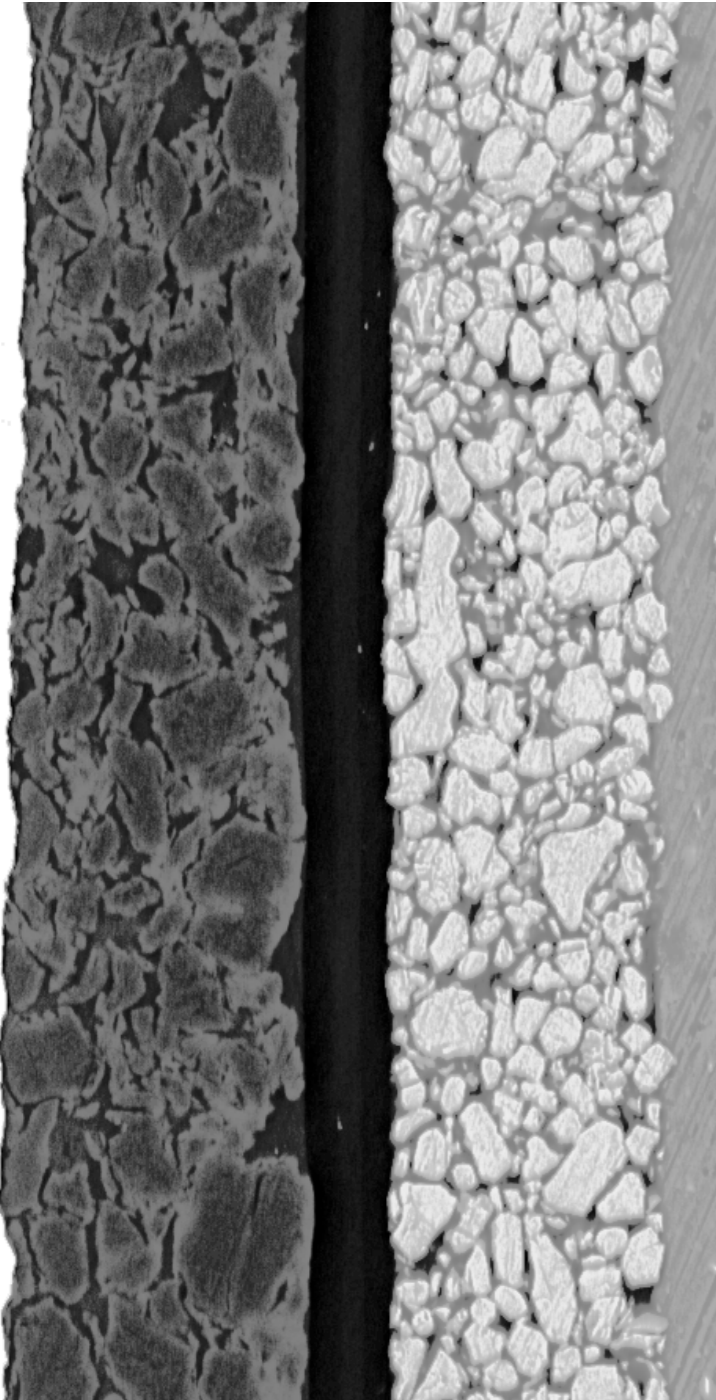
- Charge in both electrodes should be balanced, $zFC_T^C \sim zFC_T^A$
- I want $C_T^{C,A}$ to be as large as possible
- Voltage difference should be as large as possible, $\square \varphi_C \gg \varphi_A$
- Rough Figure of Merit:

$$E \sim \varphi_C \times zFC_T^C$$

Electrode Materials



Porosity in a Battery

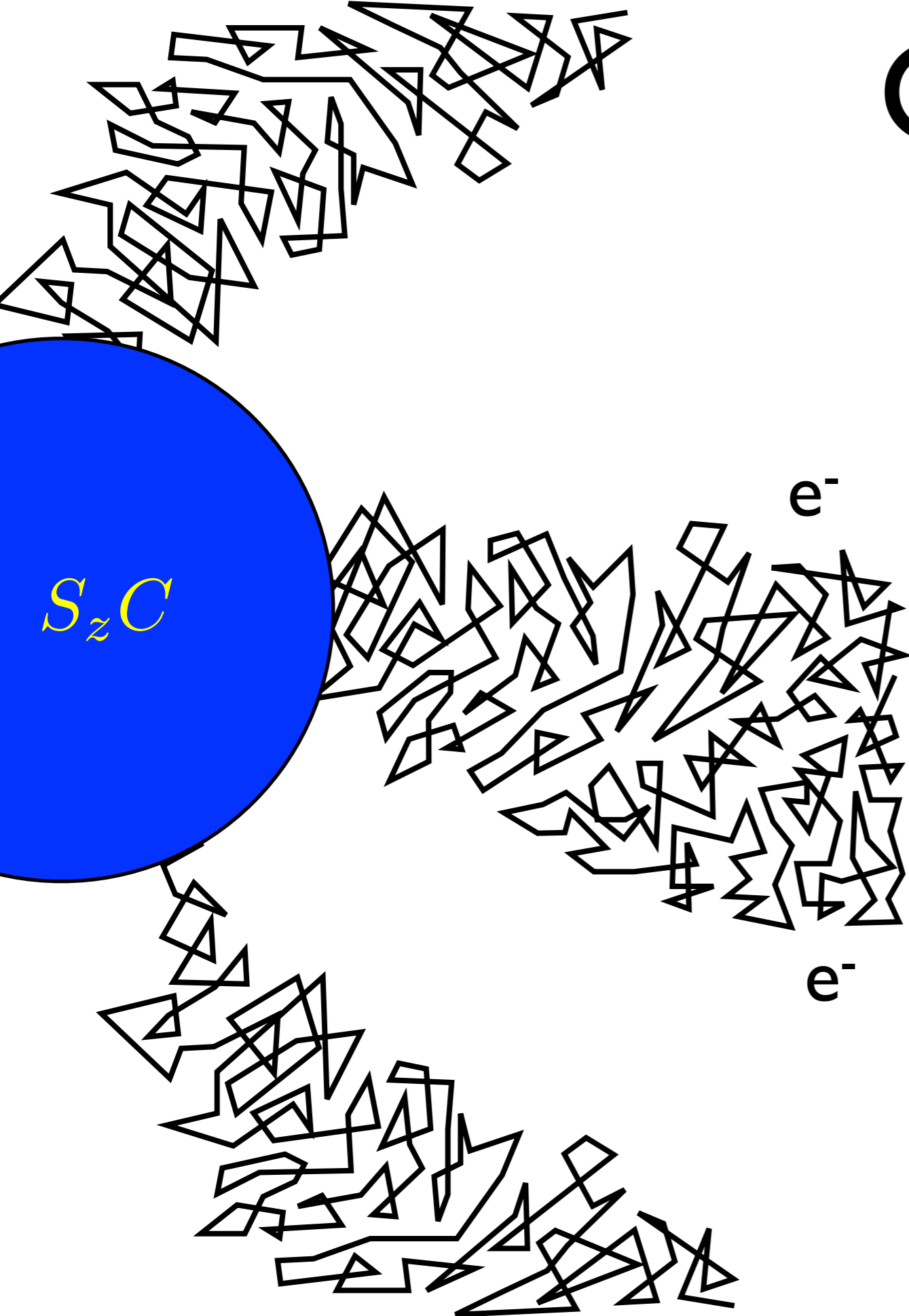


- Made porous to increase reactivity per unit volume of material
- Porosity adds additional transport paths to deliver charge faster
- Porosity is defined as volume fraction left by the active material

$$0 < \epsilon < 1$$

$$Q = z \times C_T \times V \times eN_a?$$

Other Porosity Contributions



- Ions and electrons need two routes to flow to the battery particle: need conductive carbon porosity
- Binders and fillers will occupy additional space in the electrode and will impact the capacity and energy density of cell

Porous Battery Capacity

For the Entire Device:

$$Q_c = zF c_T^c (1 - \epsilon_c - \epsilon_f) h_c$$

$$Q_a = zF c_T^a (1 - \epsilon_a - \epsilon_f) h_a$$

Capacity Ratio

Units:

mA h/m²

$$R_c = \frac{Q_c}{Q_a} = \frac{c_T^c (1 - \epsilon_c - \epsilon_f) h_c}{c_T^a (1 - \epsilon_a - \epsilon_f) h_a}$$

Units:

dimensionless

$$V = A \times h$$

